

Αντιδράσεις διπλής αντικατάστασης

Να συμπληρωθούν οι αντιδράσεις και να γραφεί η φυσική κατάσταση κάθε ένωσης

1.	$\text{AgNO}_{3(\text{aq})}$	+	$\text{H}_2\text{S}_{(\text{aq})}$	→					
2.	$\text{Na}_2\text{SO}_{4(\text{aq})}$	+	$\text{HCl}_{(\text{g})}$	→					
3.	$\text{KNO}_{3(\text{aq})}$	+	$\text{H}_2\text{S}_{(\text{g})}$	→					
4.		+		→	$\text{Ca}(\text{OH})_{2(\text{s})}$	+	$\text{KNO}_{3(\text{aq})}$		
5.	$\text{AgI}_{(\text{aq})}$	+	$\text{NH}_4\text{OH}_{(\text{aq})}$	→					
6.	$\text{NaCN}_{(\text{aq})}$	+	$\text{KOH}_{(\text{aq})}$	→					
7.		+		→	$\text{HCl}_{(\text{g})}$	+	$\text{Mg}(\text{OH})_{(\text{s})}$		
8.	$(\text{NH}_4)_2\text{SO}_{4(\text{aq})}$	+	$\text{CaCl}_{2(\text{s})}$	→					
9.	$\text{NaBr}_{(\text{aq})}$	+	$\text{Ca}(\text{OH})_{2(\text{aq})}$	→					
10.	$\text{Na}_2\text{CO}_{3(\text{s})}$	+	$\text{H}_2\text{SO}_{4(\text{aq})}$	→					
11.	$\text{NH}_4\text{I}_{(\text{aq})}$	+	$\text{Zn}(\text{OH})_{2(\text{s})}$	→					
12.	$\text{Ba}_3(\text{PO}_4)_{2(\text{s})}$	+	$\text{HBr}_{(\text{aq})}$	→					
13.	$\text{KCl}_{(\text{aq})}$	+	$\text{NaNO}_{3(\text{s})}$	→					
14.	$(\text{NH}_4)_2\text{CO}_{3(\text{aq})}$	+	$\text{Ca}(\text{OH})_{2(\text{s})}$	→					
15.	$\text{CaBr}_{2(\text{s})}$	+	$\text{NH}_4\text{I}_{(\text{aq})}$	→					
16.	$\text{AlCl}_{3(\text{s})}$	+	$\text{Mg}(\text{CN})_{2(\text{aq})}$	→					
17.		+		→	$\text{Ca}(\text{NO}_3)_{2(\text{aq})}$	+	$\text{CO}_{2(\text{g})}$	+	$\text{H}_2\text{O}_{(\text{e})}$
18.	$(\text{NH}_4)_2\text{SO}_{4(\text{aq})}$	+	$\text{KOH}_{(\text{aq})}$	→					
19.	$\text{BaCO}_{3(\text{aq})}$	+	$\text{HCl}_{(\text{g})}$	→					
20.		+		→	$\text{MgCl}_{2(\text{aq})}$	+	$\text{NH}_{3(\text{e})}$	+	$\text{H}_2\text{O}_{(\text{e})}$

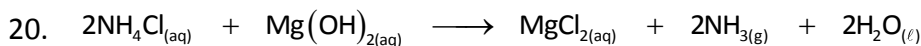
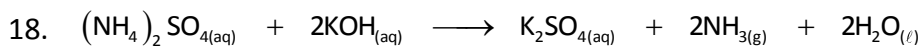
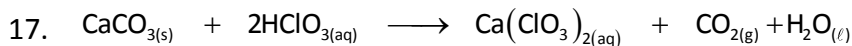
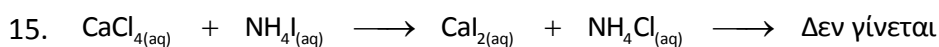
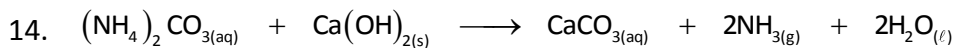
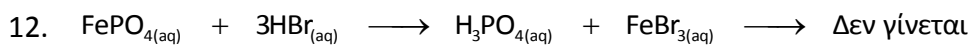
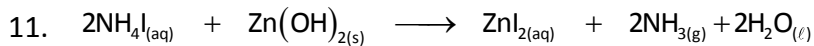
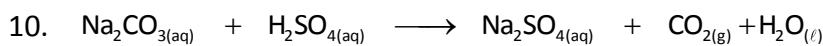
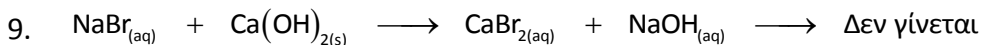
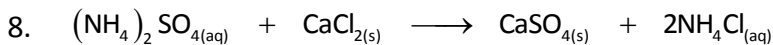
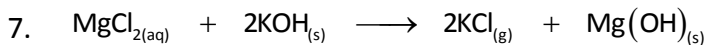
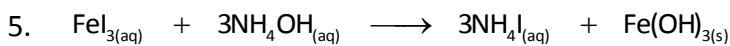
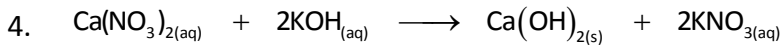
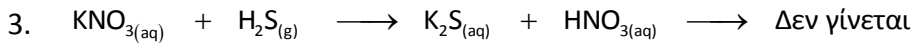
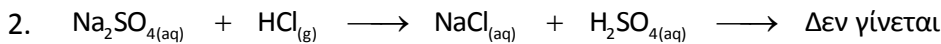
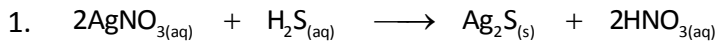
Αντιδράσεις Εξουδετέρωσης

Να συμπληρωθούν οι αντιδράσεις και να γραφεί η φυσική κατάσταση κάθε ένωσης

1.	$\text{AgOH}_{(\text{aq})}$	+	$\text{H}_2\text{S}_{(\text{aq})}$	\rightarrow			
2.	$\text{Ca}(\text{OH})_{2(\text{s})}$	+	$\text{HCl}_{(\text{aq})}$	\rightarrow			
3.	$\text{Mg}(\text{OH})_{2(\text{aq})}$	+	$\text{H}_2\text{S}_{(\text{g})}$	\rightarrow			
4.	$\text{HBr}_{(\text{aq})}$	+	$\text{KOH}_{(\text{s})}$	\rightarrow			
5.	$\text{H}_2\text{S}_{(\text{aq})}$	+	$\text{Al}(\text{OH})_{3(\text{s})}$	\rightarrow			
6.	$\text{H}_2\text{SO}_{4(\text{aq})}$	+	$\text{NaOH}_{(\text{s})}$	\rightarrow			
7.				\rightarrow	$\text{ZnCl}_{2(\text{aq})}$	+	$\text{H}_2\text{O}_{(\text{l})}$
8.	$\text{H}_3\text{PO}_{4(\text{aq})}$	+	$\text{Ba}(\text{OH})_{2(\text{s})}$	\rightarrow			
9.	$\text{HCN}_{(\text{aq})}$	+	$\text{NH}_{3(\text{g})}$	\rightarrow			
10.	$\text{Fe}(\text{OH})_{2(\text{aq})}$	+	$\text{HCl}_{(\text{g})}$	\rightarrow			
11.	$\text{Ca}(\text{OH})_{2(\text{s})}$	+	$\text{HCN}_{(\text{aq})}$	\rightarrow			
12.		+		\rightarrow	$\text{K}_3\text{PO}_{4(\text{aq})}$	+	$\text{H}_2\text{O}_{(\text{l})}$
13.	$\text{NH}_{3(\text{g})}$	+	$\text{HNO}_{3(\text{aq})}$	\rightarrow			
14.	$\text{Fe}(\text{OH})_{3(\text{aq})}$	+	$\text{HBr}_{(\text{aq})}$	\rightarrow			
15.	$\text{H}_2\text{CO}_{3(\text{aq})}$	+	$\text{Mg}(\text{OH})_{2(\text{s})}$	\rightarrow			
16.	$\text{HCN}_{(\text{aq})}$	+	$\text{Al}(\text{OH})_{3(\text{s})}$	\rightarrow			
17.	$\text{NH}_{3(\text{g})}$	+	$\text{H}_2\text{CO}_{3(\text{aq})}$	\rightarrow			
18.	$\text{H}_2\text{SO}_{4(\text{aq})}$	+	$\text{HCN}_{(\text{g})}$	\rightarrow			
19.	$\text{H}_3\text{PO}_{4(\text{aq})}$	+	$\text{Fe}(\text{OH})_{3(\text{s})}$	\rightarrow			
20.	$\text{HCl}_{(\text{aq})}$	+	$\text{Ca}(\text{OH})_{2(\text{s})}$	\rightarrow			

Αντιδράσεις διπλής αντικατάστασης

Να συμπληρωθούν οι αντιδράσεις και να γραφεί η φυσική κατάσταση κάθε ένωσης



Αντιδράσεις Εξουδετέρωσης

Να συμπληρωθούν οι αντιδράσεις και να γραφεί η φυσική κατάσταση κάθε ένωσης

